

The H positions calculated for $\text{MnSO}_4 \cdot 4\text{H}_2\text{O}$ deviate on the average by as much as 0.2 Å from the corresponding neutron-diffraction values given by Bauer (1964) for $\text{MgSO}_4 \cdot 4\text{H}_2\text{O}$, which is isostructural with $\text{MnSO}_4 \cdot 4\text{H}_2\text{O}$. The n.m.r. vectors given for $\text{MgS}_2\text{O}_3 \cdot 6\text{H}_2\text{O}$ appear to disagree with the hydrogen-bonding scheme suggested by the X-ray investigators (Nardelli, Fava & Giraldi, 1962). Both these compounds have been reexamined in this laboratory. The resonance component lines for $\text{MnSO}_4 \cdot 4\text{H}_2\text{O}$ appear to be rather wide providing no room for improvement over the measurements made by Rao. However, the n.m.r. results for $\text{MgS}_2\text{O}_3 \cdot 6\text{H}_2\text{O}$ appear to be in excellent agreement with the predictions of Nardelli *et al.* (1962).

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References

- ALEKSANDROV, K. S., LUNDIN, A. G. & MIKHAILOV, G. M. (1960). *Kristallografiya*, **5**, 84.
 BAUR, W. H. (1964). *Acta Cryst.* **17**, 863.
 BAUR, W. H. (1965). *Acta Cryst.* **19**, 909.
 BEEVERS, C. A. (1958). *Acta Cryst.* **11**, 273.
 BEEVERS, C. A. & SCHWARTZ, C. M. (1935). *Z. Kristallogr.* **91**, 157.
 CHROBAK, J. (1934). *Z. Kristallogr.* **88**, 35.
 EL SAFFAR, Z. M. (1966a). *J. Chem. Phys.* **45**, 4643.
 EL SAFFAR, Z. M. (1966b). *J. Phys. Soc. Japan*, **21**, 1844.
 EL SAFFAR, Z. M. (1967). *J. Chem. Phys.* **46**, 396.
 EL SAFFAR, Z. M. & BROWN, G. M. (1968). *Acta Cryst.* To be published.
 EL SAFFAR, Z. M. & MURTY, C. R. K. (1962). *Acta Cryst.* **15**, 285.
 GABUDA, S. P. (1962). *Doklady Akad. Nauk SSSR*, **146**, 840.
 HARPER, J. P. (1936). *Z. Kristallogr.* **95**, 266.
 HEIDE, H. G. & BOLL-DORNBERGER, K. (1955). *Acta Cryst.* **8**, 425.
 ITOH, J., KUSAKA, R., KIRIYAMA, R. & SAITO, Y. (1957). *Mem. Inst. Sci. Ind. Res.* **14**, 1.
 ITOH, J., KUSAKA, R., YAMAGATA, Y., KIRIYAMA, R. & IBAMOTO, H. (1952). *J. Chem. Phys.* **20**, 1503.
 ITOH, J., KUSAKA, R., YAMAGATA, Y., KIRIYAMA, R. & IBAMOTO, H. (1953a). *J. Phys. Soc. Japan*, **8**, 293.
 ITOH, J., KUSAKA, R., YAMAGATA, Y., KIRIYAMA, R. & IBAMOTO, H. (1953b). *Physica*, **19**, 415.
 ITOH, J., KUSAKA, R., YAMAGATA, Y., KIRIYAMA, R., IBAMOTO, H., KANDA, T. & MASUDA, Y. (1953). *J. Phys. Soc. Japan*, **8**, 287.
 JONES, D. W. & SMITH, J. A. S. (1960). *Trans. Faraday Soc.* **56**, 638.
 KARTHA, G. (1952). *Proc. Indian Acad. Sci.* **36**, 501.
 MARTINEZ, S., GARCIA-BLANCO, S. & RIVOIR, L. (1956). *Acta Cryst.* **9**, 145.
 MCGILLAVRY, C. H., DE WILDE, J. H. & BIJVOET, J. M. (1938). *Z. Kristallogr.* **100**, 212.
 MCGRATH, J. W. & SILVIDI, A. A. (1961). *J. Chem. Phys.* **34**, 322.
 MEIER, W. M. (1960). *Z. Kristallogr.* **113**, 430.
 MURTY, C. R. K. & EL SAFFAR, Z. M. (1962). *Acta Cryst.* **15**, 536.
 NARDELLI, M., FAVA, G. & GIRALDI, G. (1962). *Acta Cryst.* **15**, 227.
 PAKE, G. E. (1948). *J. Chem. Phys.* **16**, 327.
 PENFOLD, B. R. & GRIGOR, J. A. (1959). *Acta Cryst.* **12**, 850.
 RAO, C. R. (1964). *Indian J. Pure Appl. Phys.* **2**, 172.
 SCHEIN, B. J. B., LINGAFELTER, E. C. & STEWART, J. M. (1967). *J. Chem. Phys.* **47**, 5183.
 SMITH, J. A. S. & RICHARDS, R. E. (1952). *Trans. Faraday Soc.* **48**, 307.
 SPENCE, R. D. (1955). *J. Chem. Phys.* **23**, 1166.
 SPENCE, R. D. (1958). *J. Chem. Phys.* **29**, 961.
 SPENCE, R. D., AU, R. & VAN DALEN, P. A. (1964). *Physica*, **30**, 1612.
 SPENCE, R. D. & MULLER, J. (1957). *J. Chem. Phys.* **26**, 706.
 SPENCE, R. D. & NAGARAJAN, V. (1957). *Phys. Rev.* **109**, 191.
 TAYLOR, P. G. & BEEVERS, C. A. (1952). *Acta Cryst.* **5**, 341.
 TORRIE, B. H., BROWN, I. D. & PETCH, H. E. (1964). *Canad. J. Phys.* **42**, 229.
 VAN MEERSSCHE, M., DEREPE, J. M. & LOBO, P. W. (1961). *Bull. Soc. chim. France*, p. 2217.
 VAN MEERSSCHE, M., DEREPE, J. M. & LOBO, P. W. (1962). *Bull. Soc. chim. France*, p. 1035.
 VAN MEERSSCHE, M., DEREPE, J. M. & LOBO, P. W. (1963). *Acta Cryst.* **16**, 95.
 VISWESWARAMURTHY, S. (1963). *Acta Cryst.* **16**, 933.
 ZACHARIASEN, W. H. & PLETTINGER, H. A. (1963). *Acta Cryst.* **16**, 376.

Acta Cryst. (1968). **B24**, 1133

Refinement of the structure of diaquobisglycinatonickel(II), $\text{Ni}(\text{NH}_2\text{CH}_2\text{COO})_2(\text{OH})_2$. By H. C. FREEMAN and J. M. GUSS, *School of Chemistry, University of Sydney, Sydney 2006, Australia*

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The structure of diaquobisglycinatonickel(II), determined by Stosick (*J. Amer. Chem. Soc.* (1945) **67**, 365) from three projections, has been refined with use of three-dimensional data. The metal-donor bond-lengths are Ni-N(amino) 2.08 Å, Ni-O(carboxyl) 2.06 Å, Ni-O(water) 2.10 Å, and the angle N(amino)-Ni-O(carboxyl) is 81.1°.

The structure of diaquobisglycinatonickel(II) was determined by Stosick (1945) from three incompletely resolved projections. In order to obtain accurate values of bond lengths and angles for subsequent comparison with the dimensions of nickel-peptide complexes, the structure has been reinvestigated with three-dimensional data.

The cell data are shown in Table 1. The unit-cell dimensions were determined from 104 values of θ between 58° and 81°, measured on $h0l$ and $hk0$ Weissenberg photographs calibrated with aluminum powder lines. The cell parameters

were refined in ten cycles of non-linear least-squares as functions of $\sin^2 \theta$. The density was determined by flotation.

Table 1. *Crystal data*

$\text{C}_4\text{H}_{12}\text{O}_6\text{N}_2\text{Ni}$	F.W. 242.7
Monoclinic. $a = 7.625(5)$, $b = 6.596(5)$, $c = 9.670(5)$ Å,	
$\beta = 116.57(1)^\circ$, $U = 435.0$ Å ³ . $D_m = 1.86$, $D_x = 1.85$ g.cm ⁻³ ,	
$Z = 2$. $\lambda(\text{Cu } K\alpha_1) = 1.5405$ Å, $\lambda(\text{Cu } K\alpha_2) = 1.5443$ Å, $a(\text{Al}) =$	
4.04925 Å, $\mu = 34$ cm ⁻¹ . Space group $P2_1/c$ from systematic	
absences ($0k0$ absent for $k = 2n + 1$, $h0l$ absent for $l = 2n + 1$).	

Two overlapping sets of data, $h0l$ through $h5l$ and $hk0$ through $hk8$, were recorded with the same crystal specimen. $Cu K\alpha$ radiation was used. The intensities were measured visually on Weissenberg photographs, by comparison with a series of calibrated spots. The number of independent reflexions was 967, of which 74 were unobservably weak. The data were correlated by the least-squares method of Rae (1965). Lorentz, polarization and Tuell factors, but not absorption corrections, were applied.

A modified version of the program ORFLS (Busing, Martin & Levy, 1962) was used for the full-matrix least-squares refinement. The function minimized was $\sum w|F_o| - |F_c|^2$. The weights w were given by the function $w = 0.632/[1 + ((F_o - 15)/12)^2]$, derived from a plot of $|F_o| -$

$|F_c|$ versus $|F_o|$. The atomic scattering factors of Cromer & Waber (1965) for C, N, O and Ni^{2+} were used. The values for Ni^{2+} were corrected for the real part of the anomalous scattering by subtracting 3.1 electrons over the entire $\sin \theta$ range (International Tables for X-ray Crystallography, 1962). In the fifth cycle of refinement with anisotropic temperature factors for all atoms, no parameter shifted by more than 0.02 e.s.d. The final residual was $R = 0.095$. The positions of the hydrogen atoms could not be inferred with certainty from a difference Fourier synthesis computed after the refinement converged.

The final positional and thermal parameters are shown in Table 2, the observed and calculated structure amplitudes in Table 3, and the interatomic distances and bond angles

Table 2. Final positional parameters ($\times 10^4$), and coefficients B_{1j} ($\times 10^4$) in the expression $\exp - \{B_{11}h^2 + B_{22}k^2 + B_{33}l^2 + 2B_{23}kl + 2B_{13}hl + 2B_{12}hk\}$, (all with their standard deviations in parentheses)

Table with 13 columns: Ni, N, C(1), C(2), O(1), O(2), O(3), x/a, y/b, z/c, B11, B22, B33, B12, B13, B23. Rows list atoms and their coordinates and coefficients.

Table 3. Observed and calculated structure amplitudes ($\times 10$)

Large data table with multiple columns for observed (O) and calculated (C) structure amplitudes for various hkl indices. Includes headers like h k l, F_o, F_c for multiple reflections.

Table 4. Distances and angles in diaquobisglycinatonickel(II)

Code for superscripts:

Atom at			Atom at				
(none)	$x,$	$y,$	z	v	$1-x,$	$\frac{1}{2}+y,$	$\frac{1}{2}-z$
'	$-x,$	$-y,$	$-z$	vi	$-x,$	$\frac{1}{2}+y,$	$\frac{1}{2}-z$
''	$x,$	$\frac{1}{2}-y,$	$\frac{1}{2}+z$	vii	$x,$	$-1+y,$	$-z$
'''	$-x,$	$-\frac{1}{2}+y,$	$\frac{1}{2}-z$	viii	$1-x,$	$-\frac{1}{2}+y,$	$\frac{1}{2}-z$
iv	$x,$	$1+y,$	z	ix	$x,$	$\frac{1}{2}-y,$	$-\frac{1}{2}+z$

(a) Bond-lengths (E.s.d.'s=0.01 Å)		(b) Bond angles (E.s.d.'s in parentheses)	
Ni—N	2.08 Å	N—Ni—O(1)	81.1 (5)°
Ni—O(1)	2.06	N—Ni—O(3)	89.2 (5)
Ni—O(3)	2.10	O(1)—Ni—O(3)	87.8 (5)
N—C(1)	1.47	Ni—N—C(1)	108.1 (5)
C(1)—C(2)	1.52	N—C(1)—C(2)	111.1 (6)
C(2)—O(1)	1.28	C(1)—C(2)—O(1)	116.8 (7)
C(2)—O(2)	1.26	C(1)—C(2)—O(2)	118.4 (7)
		O(1)—C(2)—O(2)	124.8 (9)
		C(2)—O(1)—Ni	115.0 (6)

(c) Hydrogen-bonds and closest non-bonded contact (E.s.d.'s=0.01 Å)			
O(1)···H—O(3 ^{vi})	O(3)—H···O(1 ^{'''})	2.73 Å	
O(2)···H—O(3 ^{iv})	O(3)—H···O(2 ^{vii})	2.67	
O(2)···H—N ^{''}	N—H···O(2 ^{ix})	3.07	
O(2)·····N ^v	N·····O(2 ^{viii})	3.31	

(d) Bond angles at hydrogen-bonded atoms			
Ni—N·····O(2 ^{ix})	116.1°	C(2)—O(2)···O(3 ^{iv})	127.2°
C(1)—N·····O(2 ^{ix})	98.4	N ^{''} ·····O(2)···O(3 ^{iv})	79.0
Ni—O(1)···O(3 ^{vi})	114.8	Ni—O(3)···O(1 ⁱⁱⁱ)	122.7
C(2)—O(1)···O(3 ^{vi})	128.0	Ni—O(3)···O(2 ^{vii})	112.0
C(2)—O(2)···N ^{''}	106.3	O(1 ^{'''})···O(3)···O(2 ^{vii})	105.7

in Table 4. The oxygen atom of the water molecule is represented by O(3). Both single and stereoscopic views of the structure have already been published elsewhere (Freeman, Guss & Sinclair, 1968; Freeman, 1967b).

Discussion

Comparisons of this structure with those of glycine (Marsh, 1958) and bisglycinatocopper(II) hydrate (Freeman, Snow, Nitta & Tomita, 1964) show that the dimensions of the amino acid are not significantly changed by chelation. Because the metal-ligand bond-lengths in the Ni(II) complex are 0.08 Å longer, the N(amino)-metal-O(carboxyl) angles in the chelate rings are 4° smaller than in the Cu(II) complex. The degree of 'puckering' of the glycinatonickel(II) chelate rings is among the highest recorded for α -amino acid complexes (Freeman, 1967a). The Ni and N(amino) atoms lie at 0.26 and 0.61 Å from the plane of the carboxyl group, respectively [compared with deviations of 0.006, -0.103 Å and 0.126, 0.162 Å in the crystallographically independent chelate rings in bisglycinatocopper(II) hydrate]. The only respect in which our description of the structure differs from Stosick's is that we do not identify the vector from the N(amino) atom to O(2^{viii}) as a potential hydrogen bond. The atoms are 3.13 Å apart and the interatomic

vector makes unacceptable angles with the bonds at the N(amino) atom.

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References

- BUSING, W. R., MARTIN, K. O. & LEVY, H. A. (1962). Report ORNL-TM-305. Oak Ridge National Laboratory, Tennessee.
- CROMER, D. T. & WABER, J. T. (1965). *Acta Cryst.* **18**, 104.
- FREEMAN, H. C. (1967a). *Advanc. Protein Chem.* **22**, 303.
- FREEMAN, H. C. (1967b). *Advanc. Protein Chem.* **22**, 391.
- FREEMAN, H. C., GUSS, J. M. & SINCLAIR, R. L. (1968). *Chem. Comm.* p. 485.
- FREEMAN, H. C., SNOW, M. R., NITTA, I. & TOMITA, K. (1964). *Acta Cryst.* **17**, 1463.
- International Tables for X-ray Crystallography* (1962). Vol. III, p. 214. Birmingham: Kynoch Press.
- MARSH, R. E. (1958). *Acta Cryst.* **11**, 654.
- RAE, A. D. (1965). *Acta Cryst.* **19**, 683.
- STOSICK, A. J. (1945). *J. Amer. Chem. Soc.* **67**, 365.